General Chemistry 2 Quarter 1 Week 4

Module 4 - 5

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**Monday**

Module 4 [Properties of Water (Molecular Structure and Intermolecular Forces)]

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| PRE-TEST  1. C 2. B 3. A 4. D 5. C  ELICIT  Water  ENGAGE   1. H2O Lewis Structure - Drawing Method of H2O Lewis Structure, Molecular  Geometry of H2O, Polarity and Hybridisation in H2O molecule, with FAQs 2. Hydrogen bond   EXPLORE  1. a) The forces of attraction of water molecules is stronger than the forces of attraction exhibited by the alcohol.  b) high boiling point  2. a) Water, because it requires high energy to break the bond  b) Water. Because it also requires significant energy for the water molecules to release  c) High specific heat.  3. a) Ice is less dense in water.  b) Water has open structure due to the repulsion of the negatively-charged oxygen atoms.  4. The fluid in our body when we are sweating evaporates quickly and cool the body rapidly.  5. a) There is a force binding a liquid together and a force that binds the liquid to  another surface, just like the steam that sucks up water.  b) High surface tension | EXPLAIN  1. By increasing temperature or heat energy.  2. KE will be increased by increasing the temperature which to the escape in IMFA.  3. When the water molecules are heated enough to break its bond the additional heat energy will be shared to other water molecules and collide to others for the heat energy from a heat source be distributed.  4. By lowering the temperature.  5. The water molecules contracts.  6. Because ice has open structure due to repulsion of negative-charged oxygen atoms.  7. Water can evaporate, so, for human the body heat is used to vaporized sweat and for  plants heat is also used to convert liquid water to water vapor that goes to the atmosphere.  8. The water molecules come in contact with each other that causes for the surface layer to be created.  9. Due to the surface tension.  EVAULUATION  1. Water  2. Water  3. Increases  4. Less dense  5. Water  6. Hydrogen  7. Contracts  8. Large  9. High  10. Water  PRE-TEST  1. B 2. C 3. D 4. C 5. A |

Quarter 1- Module 5 (Crystalline and Amorphous Solids)

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| GIVE THIS A TRY:  1. A  2. A  3. B  4. C  5. A  Activity 1: 4 PIX ONE WORD:  Solid  Activity 2: Explain Further!  1. Based from the attained arrangement for both solids in the experiment; crystalline  solids have a definite internal atomic structure that follows a regular repeating pattern while  amorphous solids have irregular internal atomic structure, and as a result have a much more  swirly and irregular-looking exterior form.  2. Based on the experiment it shows that crystalline solids have a definite shape with  orderly arranged atoms (candies) while amorphous solids have a disordered array of  components not showing a definite shape. | Activity 3: Differences  1. Crystalline solids are arranged in fixed geometric patterns or the particles are  arranged in a regular repeating pattern while amorphous solids have a random orientation of  particles.  2. Crystalline solids have well-defined edges and faces while amorphous solids have  irregular or curved surfaces.  3. The structures of crystalline solids are built from repeating units called crystal lattices  while amorphous solids have random organization.  Activity 4: Classify It  1. Crystalline Solid  2. Amorphous Solid  3. Crystalline Solid  4. Crystalline Solid  5. Crystalline Solid  6. Amorphous Solid  7. Amorphous Solid  8. Amorphous Solid  9. Crystalline Solid  10. Crystalline Solid  Your Post Test  1. A  2. D  3. B  4. A  5. B |